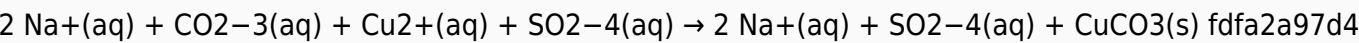


Mcgraw Hill Ryerson Chemistry 11 Solutions



Oil spill Anderson, E. These spills can have severe environmental and economic consequences. L. Howlett, K. Oil spills can result from the release of crude oil from tankers, offshore platforms, drilling rigs, and wells. Reed An oil spill is the release of a liquid petroleum hydrocarbon into the environment, especially the marine ecosystem, due to human activity, and is a form of pollution. 4th ed. 98. , E. New York: McGraw-Hill, 2003. They may also involve spills of refined petroleum products, such as gasoline and diesel fuel, as well as their by-products. Jayko, V. Kolluru, M. The term is usually given to marine oil spills, where oil is released into the ocean or coastal waters, but spills may also occur on land. Wastewater Engineering, Treatment and Reuse. Eddy. Additionally, heavier fuels used by large ships, such as bunker fuel, or spills of any oily refuse or waste oil, contribute to such incidents

For example, in the reaction of aqueous solutions of sodium carbonate and copper(II) sulfate: Oil spills penetrate into the structure of the... So...

List of group-0 ISBN publisher codes A list of publisher codes for (978) International Standard Book Numbers with a group code of zero. {{cite book}}: |first1= has generic name (help) Lewis, Elizabeth Foreman (11 May 1999). ISBN 0-8335-4007-6. Turtleback. Young Fu of the Upper Yangtze

The Na^+ and SO_4^{2-} ions are spectator ions since they remain unchanged on both sides of the equation. They simply "watch" the other ions react and does not participate in any reaction, hence the name. They are present in total ionic equations to balance the charges of the ions. Whereas the Cu^{2+} and CO_3^{2-} ions combine to form a precipitate of solid CuCO_3 . In reaction stoichiometry, spectator ions are removed from a complete ionic equation to form a net ionic equation. For the above example this yields:

Spectator ion 1. Chemistry 11 (Textbook) (1st ed. ISBN 978-0070886810. reaction. v t e A spectator ion is an ion that exists both as a reactant and a product in a chemical equation of an aqueous solution. McGraw-Hill Ryerson School. Chapter 9.). Catalysis Mustoe, Frank (May 22, 2001)

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